

Solubility Rules

All compounds containing NO_3^- , $\text{C}_2\text{H}_3\text{O}_2^-$, ClO_3^- and ClO_4^- are **Soluble** (aq).

All compounds containing an alkali metal cation* or NH_4^+ are **Soluble** (aq).

All compounds containing Cl^- , Br^- and I^- are **Soluble** (aq) except those containing Ag^+ , Hg_2^{2+} , Pb^{2+} and Cu^+ .

All compounds containing OH^- and S^{2-} are **Insoluble** (s) except those containing an alkali metal cation*, NH_4^+ , Ca^{2+} , Sr^{2+} and Ba^{2+} .

All compounds containing CO_3^{2-} and PO_4^{3-} are **Insoluble** (s) except those containing an alkali metal cation* and NH_4^+ .

All compounds containing SO_4^{2-} are **Soluble** (aq) except those containing Sr^{2+} , Ba^{2+} , Ag^{+**} , Hg_2^{2+} and Pb^{2+} .

*Alkali metal cations are Li^+ , Na^+ , K^+ , Rb^+ , and Cs^+

** Ag_2SO_4 is slightly soluble and is sometimes listed as soluble